Chemical Calculations Practice Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Chapter 7 Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

1. What is the molar mass of CaSO4?
2. Calculate the mass of 2.50 mol iron (II) hydroxide.
3. Calculate the moles in 75.0 grams dinitrogen trioxide.
4. What volume does 0.335 mol C2H6 have at STP?
5. What is the percent composition when 9.03g Mg combines completely with 3.48g N?
6. What is the percent composition of C2H6?
7. What is the percent composition of NH3?
8. What is the empirical formula for a compound which is

67.6% Hg, 10.8% S, and 21.6% O?

1. Find the molecular formula of ethylene glycol? Its empirical formula is CH3O and its molar mass is 62 g/mol.